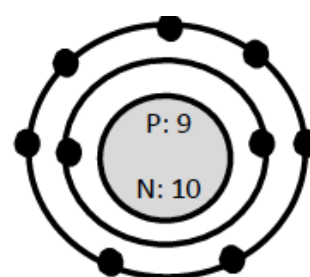
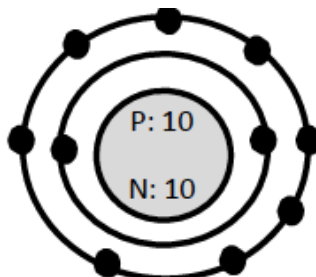
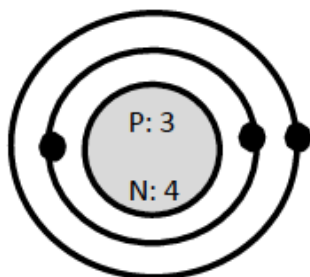


Name: _____ Date: _____ Period: _____

Periodic Table Review

Directions: Complete the chart using a periodic table.

M- metals NM- Nonmetals MT- metalloids



Element			
M, NM, MT			
Reactivity			

Directions: fill in the blanks to the statements below.

Vertical columns on the periodic table are called _____.

Horizontal rows on the periodic table are called _____.

The number of protons in an atom is that element's _____ number.

The total of the protons and neutrons is that atom's _____ number.

The elements in group _____ are the most reactive metals.

The elements in group _____ are the most reactive nonmetals.

The elements in group _____ are very unreactive.

Directions: Complete the chart using a periodic table.

	Element Symbol	M, NM, or MT	Valence Electrons	Reactivity (high, medium, non)
group 14, period 2				
group 2, period 2				
group 1, period 1				
group 18, period 3				
group 13, period 5				
group 17, period 4				
group 1, period 3				

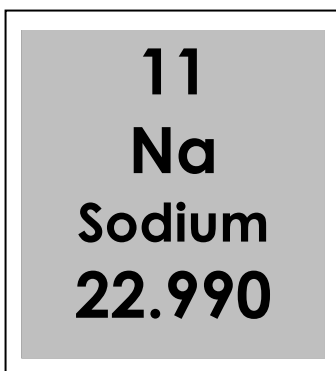
Directions: For each of the following, label as a metal (M), nonmetal (NM), and/or metalloid (MT)

Poor conductor of electricity		Silicon	
Usually a solid at room temp		Most are a gas at room temp.	
Ductile		Cobalt	
Chlorine		Good Conductor	
Semi-conductor		Brittle	
Malleable		Oxygen	

List the 3 subatomic particles in an atom.

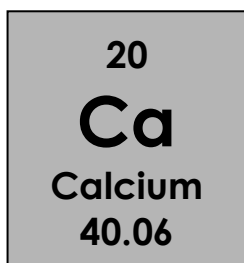
Atomic Number is equal to...	Atomic Mass is equal to...

Label the parts in the periodic square of sodium.



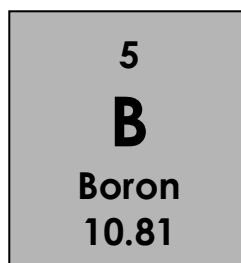
A		M	
P		A	
E		N	

Directions: Complete APEMAN for the elements below



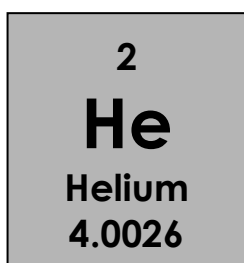
A=
P=
E=

M=
A=
N=



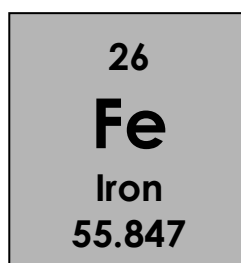
A=
P=
E=

M=
A=
N=




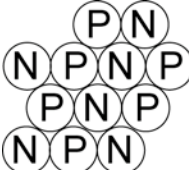

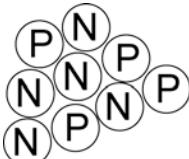
A=
P=
E=

M=
A=
N=



A=
P=
E=

M=
A=
N=

Nucleus of Atom	Element	Electrons	Atomic Mass
			
			
			
			

Directions: label the trends of the periodic table by adding arrows and descriptions.

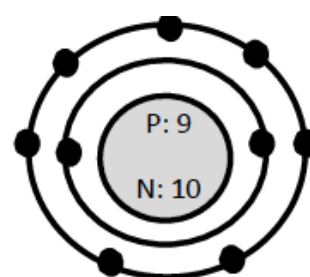
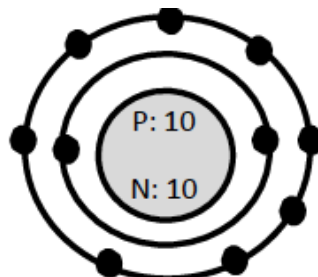
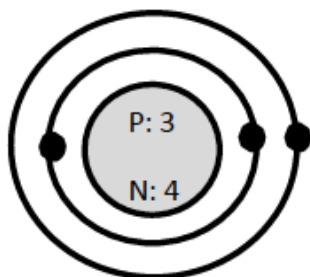
1 1A	2 2A	3 3B	4 4B	5 5B	6 6B	7 7B	8	9	10	11 1B	12 2B	13 3A	14 4A	15 5A	16 6A	17 7A	18 8A
1 H 1.008 Hydrogen																	2 He 4.003 Helium
2 3 Li 6.941 Lithium	4 Be 9.012 Beryllium											5 B 10.812 Boron	6 C 12.011 Carbon	7 N 14.007 Nitrogen	8 O 15.999 Oxygen	9 F 18.998 Fluorine	10 Ne 20.180 Neon
3 11 Na 22.990 Sodium	12 Mg 24.305 Magnesium											13 Al 26.982 Aluminum	14 Si 28.086 Silicon	15 P 30.974 Phosphorus	16 S 32.066 Sulfur	17 Cl 35.453 Chlorine	18 Ar 39.948 Argon
4 19 K 39.098 Potassium	20 Ca 40.078 Calcium	21 Sc 44.956 Scandium	22 Ti 47.867 Titanium	23 V 50.942 Vanadium	24 Cr 51.996 Chromium	25 Mn 54.938 Manganese	26 Fe 55.845 Iron	27 Co 58.933 Cobalt	28 Ni 58.693 Nickel	29 Cu 63.546 Copper	30 Zn 65.38 Zinc	31 Ga 69.723 Gallium	32 Ge 72.64 Germanium	33 As 74.922 Arsenic	34 Se 78.96 Selenium	35 Br 79.904 Bromine	36 Kr 83.798 Krypton
5 37 Rb 85.468 Rubidium	38 Sr 87.62 Strontium	39 Y 88.906 Yttrium	40 Zr 91.224 Zirconium	41 Nb 92.906 Niobium	42 Mo 95.94 Molybdenum	43 Tc 98 Technetium	44 Ru 101.07 Ruthenium	45 Rh 102.906 Rhodium	46 Pd 106.42 Palladium	47 Ag 107.868 Silver	48 Cd 112.412 Cadmium	49 In 114.818 Indium	50 Sn 118.711 Tin	51 Sb 121.760 Antimony	52 Te 127.60 Tellurium	53 I 126.905 Iodine	54 Xe 131.294 Xenon
6 55 Cs 132.905 Cesium	56 Ba 137.328 Barium	71 Lu 174.967 Lutetium	72 Hf 178.49 Hafnium	73 Ta 180.948 Tantalum	74 W 183.84 Tungsten	75 Re 186.207 Rhenium	76 Os 190.23 Osmium	77 Ir 192.225 Iridium	78 Pt 195.085 Platinum	79 Au 196.967 Gold	80 Hg 200.59 Mercury	81 Tl 204.383 Thallium	82 Pb 207.2 Lead	83 Bi 208.980 Bismuth	84 Po [209] Polonium	85 At [210] Astatine	86 Rn [222] Radon
7 87 Fr [223] Francium	88 Ra [226] Radium	103 Lr [260] Lawrencium	104 Rf [261] Rutherfordium	105 Db [262] Dubnium	106 Sg [263] Seaborgium	107 Bh [264] Bohrium	108 Hs [265] Hassium	109 Mt [266] Meitnerium	110 Ds [267] Darmstadtium	111 Rg [268] Roentgenium	Mass numbers in parentheses are those of the most stable or most common isotopes.						
Lanthanide Series		57 La 138.905 Lanthanum	58 Ce 140.116 Cerium	59 Pr 140.908 Praseodymium	60 Nd 144.242 Neodymium	61 Pm [145] Promethium	62 Sm 150.36 Samarium	63 Eu 151.964 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.925 Terbium	66 Dy 162.500 Dysprosium	67 Ho 164.930 Holmium	68 Er 167.259 Erbium	69 Tm 168.934 Thulium	70 Yb 173.055 Ytterbium		
Actinide Series		89 Ac [227] Actinium	90 Th 232.038 Thorium	91 Pa 231.036 Protactinium	92 U 238.029 Uranium	93 Np [237] Neptunium	94 Pu [244] Plutonium	95 Am [243] Americium	96 Cm [247] Curium	97 Bk [247] Berkelium	98 Cf [251] Californium	99 Es [252] Einsteinium	100 Fm [257] Fermium	101 Md [258] Mendelevium	102 No [259] Nobelium		

Name: _____ Date: _____ Period: _____

Periodic Table Review **Answer Key**

Directions: Complete the chart using a periodic table.

M- metals NM- Nonmetals MT- metalloids



Element	lithium	neon	Fluorine
M, NM, MT	M	NM	NM
Reactivity	Highly Reactive	Non-Reactive	Highly Reactive

Directions: fill in the blanks to the statements below.

Vertical columns on the periodic table are called groups.

Horizontal rows on the periodic table are called periods.

The number of protons in an atom is that element's atomic number.

The total of the protons and neutrons is that atom's mass number.

The elements in group 1 are the most reactive metals.

The elements in group 17 are the most reactive nonmetals.

The elements in group 18 are very unreactive.

Directions: Complete the chart using a periodic table.

	Element Symbol	M, NM, or MT	Valence Electrons	Reactivity (high, medium, non)
group 14, period 2	Si	MT	4	medium
group 2, period 2	Be	M	2	medium
group 1, period 1	H	NM	1	high
group 18, period 3	Ar	NM	8	non
group 13, period 5	In	M	3	medium
group 17, period 4	Br	NM	7	high
group 1, period 3	Na	M	1	high

Directions: For each of the following, label as a metal (M), nonmetal (NM), and/or metalloid (MT)

Poor conductor of electricity	NM	Silicon	MT
Usually a solid at room temp	M	Most are a gas at room temp.	NM
Ductile	M	Cobalt	M
Chlorine	NM	Good Conductor	M
Semi-conductor	MT	Brittle	NM, MT
Malleable	M	Oxygen	NM

List the 3 subatomic particles in an atom.

Protons

Neutrons

Electrons

Atomic Number is equal to...	Atomic Mass is equal to...
Protons	(nucleus of the atom) Protons + Neutrons

Label the parts in the periodic square of sodium.

11
Na
Sodium
22.990

Atomic Number

Symbol

Element Name

Mass Number

A	Atomic Number] Equal	M	Mass Number
P	Protons		A	Atomic Number (Subtract)
E	Electrons		N	Neutrons

Directions: Complete APEMAN for the elements below

20
Ca
Calcium
40.06

A= 20 M= 40
P= 20 A= 20
E= 20 N= 20

5
B
Boron
10.81

A= 5 M= 11
P= 5 A= 5
E= 5 N= 6

2
He
Helium
4.0026

A= 2 M= 4
P= 2 A= 2
E= 2 N= 2

26
Fe
Iron
55.847

A= 26 M= 56
P= 26 A= 26
E= 26 N= 30

Nucleus of Atom	Element	Electrons	Atomic Mass
	Li	Total= 3 Valence= 1	P= 3 N= 4 Mass= 7
	C	Total= 6 Valence= 4	P= 6 N= 6 Mass= 12
	H	Total= 1 Valence= 1	P= 1 N= 0 Mass= 1
	Be	Total= 4 Valence= 2	P= 4 N= 5 Mass= 9

Directions: label the trends of the periodic table by adding arrows and descriptions.

