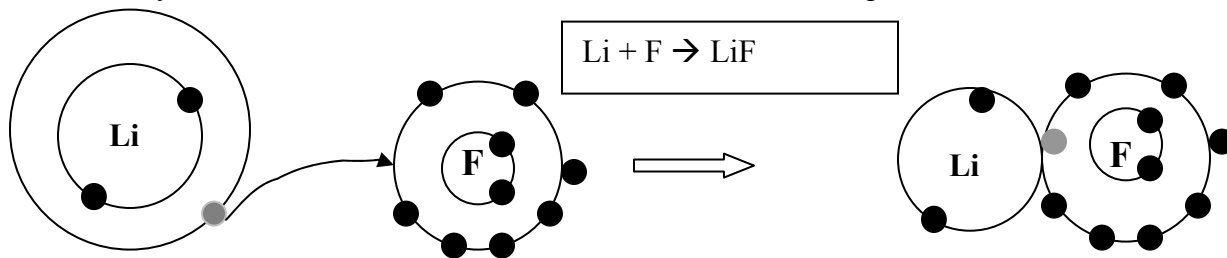


Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li			
F			

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$				
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$				
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$				
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$				
5) $\text{K} + \text{F} \Rightarrow \text{KF}$				

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
6) $\text{Al} + \text{Cl} \Rightarrow \text{AlCl}_3$				
7) $\text{Na} + \text{O} \Rightarrow \text{Na}_2\text{O}$				
8) $\text{Li} + \text{N} =$				
9) $\text{Mg} + \text{F} =$				
10) $\text{Na} + \text{F} \Rightarrow$				
11) $\text{Al} + \text{O} \Rightarrow$				
12) $\text{Li} + \text{O} \Rightarrow$				
13) $\text{K} + \text{S} \Rightarrow$				
14) $\text{Mg} + \text{O} \Rightarrow$				